NAME

DATE

CLASS

CHAPTER REVIEW

Atomic Structure and the Periodic Table

Part A. Vocabulary Review
On the space at the left, write the term that correctly completes each statement. Use each term once.

metals groups electrons chemical symbol	isotopes metalloids nucleus quarks	average atomic mass transition elements mass number periodic table	electron cloud atomic number periods							
	 A capital letter or a combination of a capital letter and a small letter that is used to represent an element is called a(n) 									
	2. The horizontal rows of elements are called									
	3. An average of the masses of all the isotopes that occur in nature for an element is the									
	4. Vertical columns of elements are called									
	5. Elements in th are called the	e middle of the periodic tab	le, periods 4 through 7,							
	6. The number of	f protons in an atom is the _								
•	7. Protons and ne them.	eutrons can be subdivided in	nto by colliding							
	8. The center of a	n atom where protons and 1	neutrons are located is the							
	9. A total count o	f the neutrons and protons i	n an atom is the							
	10. Atoms of the sa are	ame element but with differ	ent numbers of neutrons							
-	11. Elements that a	re found on the left side of	the periodic table are							
	12. Elements that h are	ave some properties of both	n metals and nonmetals							
	13. The particles the	at move about the nucleus a	and have a negative							
	14. The region arou	nd the nucleus occupied by	electrons is a(n)							
	15. A chart that sho	ws the classification of elem	ents is called the							

Part B. Concept Review
Complete the table below by writing the correct information in the blank spaces.

Element	Chemical Symbol	Atomic Number	Number of Protons	Number of Neutrons	Mass Number
1. Sodium		11	11	12	23
2. Carbon	С		6	6	12
3.	Fe	26	26	30	56
4. Sulfur	S	16		16	32
5. Nitrogen	N	7	7		14
6. Oxygen	0	8	8	8	
7.	He	2	2	2	4
8. Chlorine-35	Cl	17	17		35
9. Copper	Cu	29	29	35	
10. Chlorine-37	CI		17	20	37

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Refer to the periodic table below and the boxes at the right of the table to answer questions 11–15.

1	_																18
T	2											13	14	15	16	17	
	Ē	1															
-		3	4	5	6	7	8	9	10	11	12		D			С	
A		191					Fe		-								
					. "				1	Au							
В	П																

Na 23 26 55.8

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 11. Which element has a greater atomic mass, A or	B?

- ____ 12. Which element is a metal, B or C?
- ____ 13. In which group is gold?
- ____ 14. What is the average atomic mass of iron?
- ___ 15. What is the atomic number of sodium?